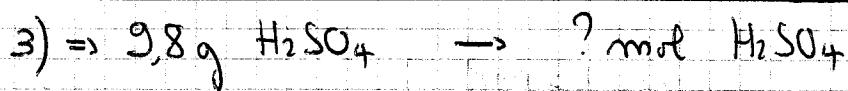
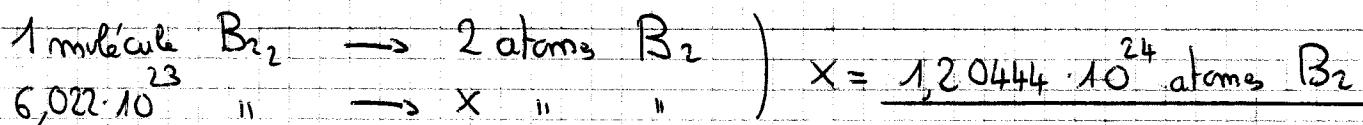


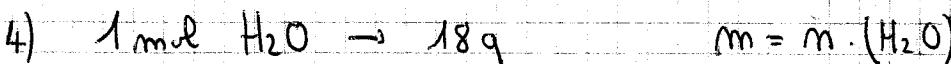
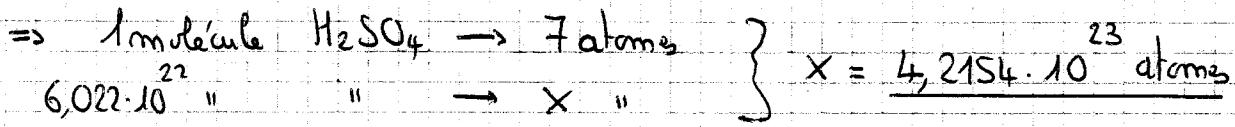
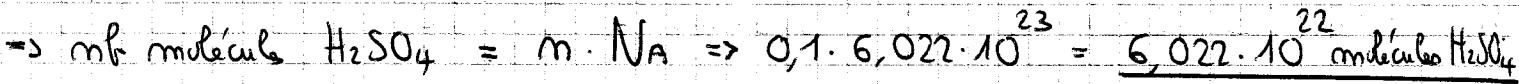
# Longes exercices "molé" (1/2)

①

- 1) 1 mol Br<sub>2</sub> pèse 159,80g
- 1 mol Al pèse 26,98g
- 1 mol Hg pèse 200,59g
- 1 mol Cu pèse 63,55g
- 1 mol S pèse 32,06g
- 1 mol Zn pèse 65,38g
- 1 mol Fe pèse 55,85g

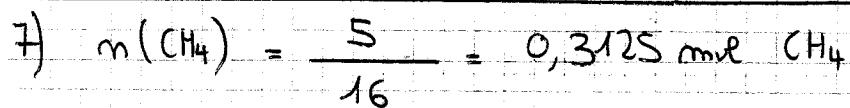
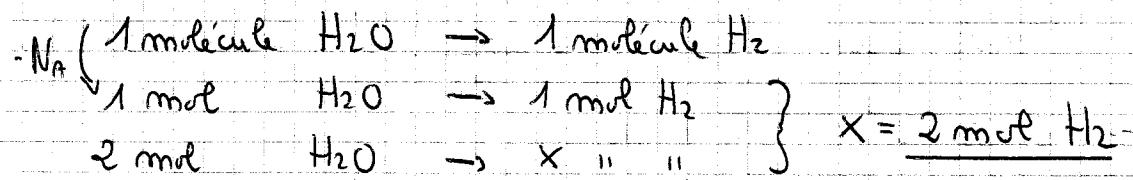


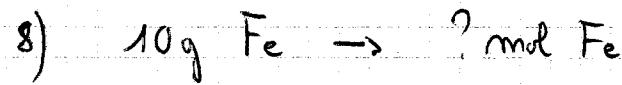
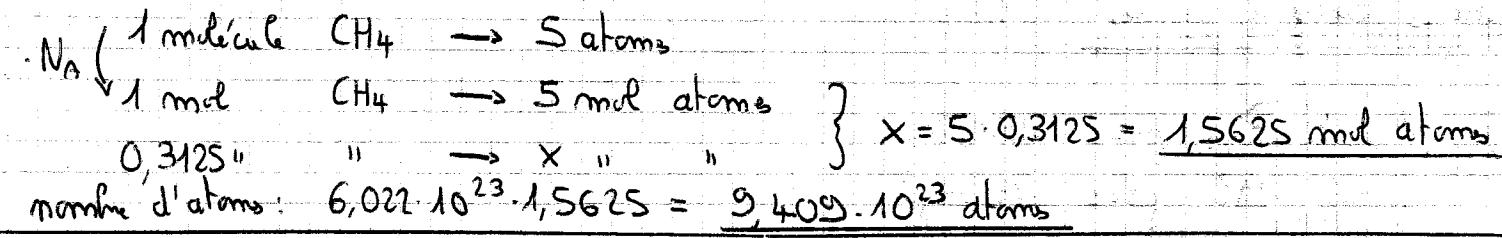
$$m(H_2SO_4) = \frac{9,8}{98} = 0,1 \text{ mol}$$



$$m = m \cdot (H_2O)$$

$$m = 1 \cdot 18 = \underline{\underline{18 \text{ g}}}$$

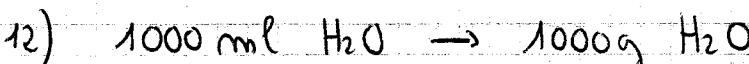
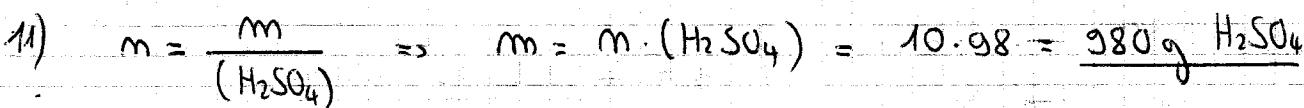
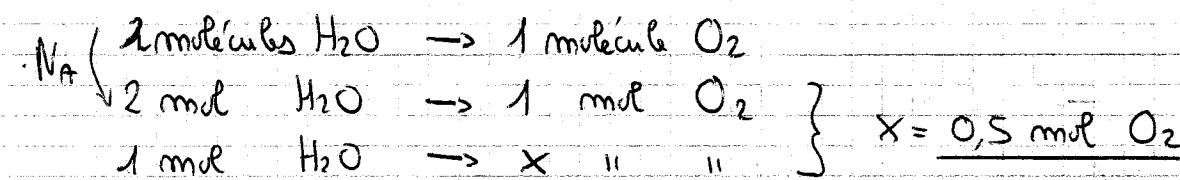
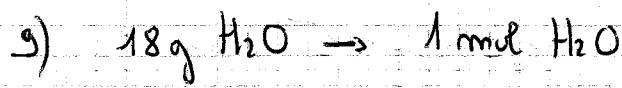




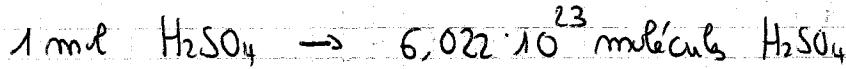
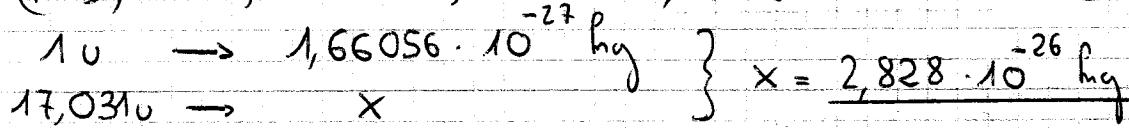
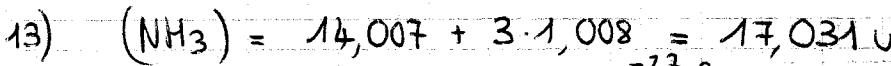
$$m = \frac{10}{55,85} = 0,17905 \text{ mol Fe}$$

$\Rightarrow$  pour avoir le même nombre d'atomes, il faut prendre le même nombre de mol!

$$\Rightarrow m(Mg) = m \cdot (Mg) = 0,17905 \cdot 24,31 = 4,3527 \text{ g Mg}$$

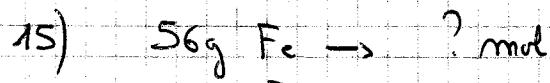


$$m = \frac{1000}{18} = 55,56 \text{ mol H}_2\text{O}$$

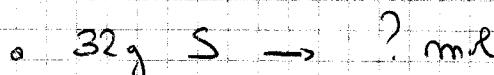


$$\Rightarrow x = 0,5102 \cdot 6,022 \cdot 10^{23} = 3,0724 \cdot 10^{23} \text{ molécules H}_2SO_4$$

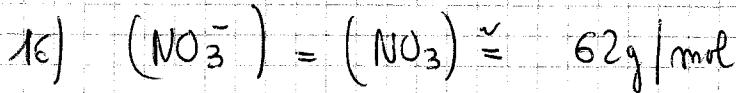
Lange's exercises "mole" (suite) (2/2)



$$\circ m = \frac{56}{55,85} = 1,003 \text{ mol Fe}$$



$$m = \frac{32}{32,06} = 0,998 \text{ mol S}$$



$$M = \frac{m}{M} \Rightarrow m = \frac{186}{62} = 3 \text{ mol } \text{NO}_3^-$$